

REMARKS

An early examination of this divisional application is respectfully requested. An INFORMATION DISCLOSURE STATEMENT is filed concurrently.

Original Claims 1-7 are canceled herein. The Claims 8-20 presented herein correspond to the claims 3-7 in the parent case. Claims 3-7 of the parent application were withdrawn from consideration, in view of a restriction requirement, in response to which applicants elected claims 1 and 2, with consequent withdrawal of Claims 3-7. Specific recitations in certain claims are noted: Claim 15 recitations are supported by page 5 of the specification, four lines from the bottom; and Claim 16 recitations are supported by page 9

A SUBSTITUTE SPECIFICATION, clean copy and marked up copy, are attached hereto. These papers are presented to conform the text of the instant application to that of the parent case when it is issued. Specifically, the same SUBSTITUTE SPECIFICATION was first filed on December 31, 2002, in the parent application to clarify a translation issue relating to the word crystal. Applicants changed the word "crystal" to --quartz -- in the specification and claims of the parent case, as explained by the DECLARATION. Applicants wish to conform the language herein to that of the present language of the parent case. The following are the remarks previously presented in the parent case concerning the change:

"Applicants enclose herewith two attachments including a clean copy and a marked up version of a substitute specification. This presentation is to conform the specification to the drawings. In

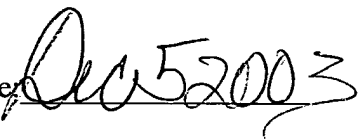
the drawings the word quartz is used; please see Figure 6.

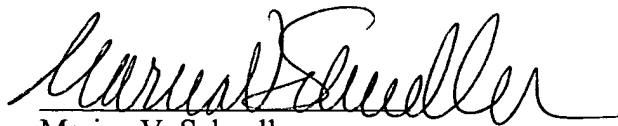
However in translation of the Japanese language specification, the Japanese word for “quartz” may also refer to crystal. This is evidenced by comparing the terms on Figure 6 referring to quartz to the terms in the specification describing Figure 6 which refers to employs the word ‘crystal’ Example 2, at page 15 line 8 thereof. Amorphous silicon dioxide and single crystalline silicon dioxide (quartz) can be represented by the same chemical formula “SiO₂”. However, single crystalline silicon dioxide has completely different physical characteristics from that of amorphous silicon dioxide. Amorphous silicon dioxide has only insulating properties and amorphous silicon dioxide thin film is often used as an insulation film for semiconductor devices. On the other hand single crystalline silicon dioxide has resonance an piezo electric characteristics. "

An early examination is respectfully solicited.

Respectfully submitted,

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QUARTZ THIN FILM

BACKGROUND OF THE INVENTION

1. Field of the Invention

The present invention relates to a quartz thin film for use in vibrators, oscillators, high frequency filter surface acoustic wave elements, optical waveguides, semiconductor substrates and the like, and to a production method therefor.

2. Description of the Related Art

Quartz thin films have been widely used for vibrators, oscillators, high frequency filter surface acoustic wave elements, optical waveguides and the like, and are very important industrial materials. Quartz thin films are produced by a method in which a single crystal of quartz obtained by a hydrothermal synthesis is polished to provide a thin film, or by a direct production method such as a sol-gel method, a plasma enhanced chemical vapor deposition (CVD) method, a sputtering method, and a laser ablation method.

Many thin film techniques for single crystals of quartz have been developed, as disclosed in Japanese Unexamined Patent Applications, Publication Nos. 8-157297, 8-225398, 8-268718, 8-91977, and 9-315897.

In these patent publications, a single crystal of quartz thin film is produced by a sol-gel method, a plasma CVD method, or a sputtering method using a silicon alkoxide as a source material. The sol-gel method requires many complex steps such as adding alcohol, water, and amine to a source solution, refluxing, coating on the substrate, drying, and heat processing. The plasma CVD method, the sputtering method, and the laser ablation method require large-scale

equipment. For example, in the laser ablation method, a sintered target is irradiated with an ultraviolet laser pulse under an ultra-high vacuum to vaporize it and produce a plasma, thereby growing a thin film on a substrate. It requires a vacuum device capable of reducing pressure to an extreme degree, an excimer laser, and a sintered material. The growth rate of the quartz thin film formed on the substrate is as low as $0.25 \mu\text{m/h}$.

In summary, the sol-gel method requires many steps and is therefore unsuitable for industrial production. The laser ablation method requires expensive equipment, such as vacuum devices, and expensive materials such as highly purified silicon dioxide as a target, and can grow a quartz thin film only at a low rate, which makes it unsuitable for industrial production.

SUMMARY OF THE INVENTION

An object of the present invention is to provide a quartz thin film for use in vibrators, oscillators, high frequency filter surface acoustic wave elements, optical waveguides, semiconductor substrates, and the like, using an atmospheric pressure vapor phase epitaxial (AP-VPE) method in which ethyl silicate is reacted with oxygen to grow a thin film epitaxially on a substrate at atmospheric pressure rather than under a vacuum.

The present inventors have discovered a method for producing a quartz thin film using the AP-VPE method in which an inexpensive silicon alkoxide is reacted with oxygen to grow a thin film epitaxially on a substrate under atmospheric pressure rather than under a vacuum.

The method for producing a quartz epitaxial thin film according to the present invention comprises the steps of vaporizing a silicon alkoxide at atmospheric pressure to introduce the silicon alkoxide on a substrate, and reacting the silicon alkoxide with oxygen to deposit a quartz on the substrate.

In the method for producing the quartz epitaxial thin film, a reaction promoter, e.g.,

hydrogen chloride, for promoting the reaction of the silicon alkoxide with oxygen, may be added. The substrate may have a buffer layer on which the quartz epitaxial thin film is grown.

According to the present invention, the deposition thickness of the quartz epitaxial thin film on the substrate is as high as about 3 μm per hour.

Examples of the silicon alkoxide for producing a quartz by reacting with oxygen include tetramethoxysilane, tetraethoxysilane, tetrapropoxysilane and tetrabutoxysilane. Examples of an oxygen-containing gas for producing a quartz by reacting with the silicon alkoxide include oxygen, ozone, dinitride monoxide, and water.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 is a schematic diagram of an apparatus for epitaxial growth of a quartz thin film according to the present invention.

FIG. 2 is a graph showing a double crystal X-ray diffraction (DC-XRD) spectrum of a quartz epitaxial growth film made with a tetraethoxysilane-oxygen-hydrogen chloride-nitrogen system.

FIG. 3 is a graph showing the relationship between an epitaxial growth temperature and a full width at half maximum (FWHM).

FIG. 4 is a graph showing the relationship between growing temperature and epitaxial growth rate of an epitaxial film.

FIG. 5 is a graph showing an infrared light absorption spectrum of a quartz thin film made with tetraethoxysilane-oxygen-hydrogen chloride-nitrogen system.

FIG. 6 is a graph showing the relationship between thickness of a buffer layer and a full width at half maximum (FWHM).

Description of the Preferred Embodiments

According to the present invention, the quartz thin film grows epitaxially by heating and vaporizing the silicon alkoxide at atmospheric pressure, which is transferred to a substrate using a carrier gas such as nitrogen and is reacted with oxygen gas on the substrate. A heating source for the silicon alkoxide is, for example, a high frequency induction heater, or a resistance heater.

The silicon alkoxide as the source material is heated and a part thereof is introduced into a growing area for growing on the substrate as a vapor phase. The amount of the silicon alkoxide introduced can be adjusted by changing a heating temperature and the amount of the carrier gas. The heating temperature ranges generally from about 50 to 120°C in view of a vapor pressure of the silicon alkoxide.

Any inert gases can be used as the carrier gas for introducing the part of the source material into the growing area. For example, nitrogen, argon, and helium can be used as the carrier gas. Nitrogen is preferred because it is inexpensive.

The silicon material does not necessarily have a purity as high as that of a target, such as 3N and 5N, which is used in a laser ablation method. It is sufficient that it have a purity of about 99.5%. The higher the purity, the more the crystallinity and the crystal properties are enhanced.

The oxygen-containing gas for producing a quartz by reacting with the silicon alkoxide is introduced into the growing area with the carrier gas. A larger amount of the oxygen-containing gas tends to increase the growth rate of the quartz thin film and to improve the crystallinity thereof. Conversely, a smaller amount of the oxygen-containing gas tends to decrease the growth rate and to degrade the crystallinity thereof. This may be caused by a small equilibrium constant of the reaction by which the silicon alkoxide and oxygen produce the quartz. It is therefore preferred that oxygen be provided in excess relative to the silicon alkoxide. The oxygen partial pressure is generally about 0.1 to 0.3 atm, depending on the amount of the silicon alkoxide

vaporized. The growth rate of the quartz increases by adding hydrogen chloride (HCL) to the reaction of the silicon alkoxide with oxygen. It is believed that hydrogen chloride functions catalytically in the reaction of the silicon alkoxide with oxygen. In this case, the oxygen partial pressure is generally about 0.001 to 0.003 atm depending on the amount of the silicon alkoxide vaporized.

Examples of the substrate for use in the present invention include a sapphire substrate, a silicon substrate, and a GaAs substrate. The substrate is heated and is maintained at a constant temperature. The substrate may be parallel to or vertical with respect to the flow of the source gas, or it may be tilted at a predetermined angle.

A quartz epitaxial thin film may be grown directly on the substrate. Alternatively, the quartz epitaxial thin film may be grown on a buffer layer formed on the substrate. This leads to improved crystallicity in the quartz epitaxial thin film.

The buffer layer can be formed by, for example, growing an amorphous crystal on the substrate at a relatively low temperature and annealing it. Alternatively, the buffer layer may be formed by pre-forming other materials such as GaN and ZnO which have a hexagonal shape similar to the quartz on the sapphire substrate. The buffer layer contributes to correcting nonconformity in the lattices.

To verify the buffer layer contribution, a quartz epitaxial thin film was formed on a sapphire substrate on which an amorphous crystal buffer layer having a thickness of 50 nm was formed, which was formed by depositing a quartz thereon at 550°C and annealing it. A full width at half maximum (FWHM) determined by an X-ray diffraction peak was 1.0 minute. Typically, the FWHM is an index of crystallicity in the epitaxial film. Another quartz epitaxial thin film was formed on a sapphire substrate with no buffer layer. The FWHM was about 10.0 minutes. The result shows that the buffer layer contributes to enhancing crystallicity in the quartz

epitaxial thin film. The smaller the FWHM, the sharper peak and the better the crystallinity.

It is believed that the crystallinity is improved because the buffer layer corrects nonconformity in lattices (1), prevents cracking caused by differential thermal expansion (2), and promotes growth in a width direction (substrate direction) in an initial growing stage (3). It is preferred that the thickness of the buffer layer be about 25 to 80 nm based on experimental results.

One example of the quartz epitaxial thin film growth using nitrogen as the carrier gas according to the present invention is explained below. A vapor phase epitaxial growing device used in the present invention is a vertical quartz reactor as shown in FIG. 1. The reactor 1 comprises a source material providing portion 2 and a growing area 3, and both are maintained at a certain temperature. The source material providing portion 2 is heated to a temperature within a range of 50 to 120°C. In FIG. 1, the source material providing portion 2 is heated to 70°C. A silicon alkoxide as the silicon source material is vaporized at a bubbler (vaporizer) 4, and is transported to the growing area 3 with nitrogen provided from the left side. Together with the nitrogen, oxygen for forming a quartz and hydrogen chloride (catalyst) for promoting the reaction are introduced into the growing area 3 in the reactor at a predetermined partial pressure.

The growing area 3 is heated to a temperature within a range of 550 to 850°C at which the quartz grows. A sapphire substrate 5 is maintained at the same temperature. Ethyl silicate is effectively reacted with oxygen by use of catalytic hydrogen chloride to produce a quartz. The quartz adsorbs, grows, and forms an epitaxial film on the sapphire (0001) substrate. The nitrogen introduced at a lower portion of the reactor is for (1) creating a convection at the growing area 3 to promote the reaction, and (2) drawing the gas toward a predetermined exhaust vent. The total pressure in the reactor is maintained at atmospheric pressure. Ethane, carbon dioxide and water that are produced upon the reaction, and unreacted tetraethoxysilane and oxygen are exhausted

together with the carrier gas, nitrogen, from the exhaust vent. In FIG. 1, a rod 6 made of quartz is for mounting, inserting, and withdrawing the sapphire substrate. TABLE 1 shows typical reaction conditions. In TABLE 1, "sccm" means "standard cubic centimeter per minute".

TABLE 1

Tetraethoxysilane partial pressure	3.3×10^{-3} atm
Oxygen partial pressure	3.3×10^{-1} atm
Hydrogen chloride	1.67×10^{-3} atm
Total pressure	1 atm
Total flow rate	800 sccm
Substrate	Sapphire (0001)
Growth temperature	550 to 850°C

The growth rate of the epitaxial film increased as the temperature increased. The maximum growth rate was 3 $\mu\text{m/h}$. This growth rate is about 12 times higher than that, i.e., about 0.25 $\mu\text{m/h}$, in a normal laser ablation method.

The quartz epitaxial thin film thus obtained by the present invention did not contain impurities and had high crystallinity. The thin film can be used in vibrators, oscillators, high frequency filter surface acoustic wave elements, optical waveguides, semiconductor substrates, and the like.

The present invention will be illustrated by the following Examples.

Example 1

The apparatus shown in FIG. 1 was used. Oxygen having a purity of 99.9% or more flowed at a flow rate of 200 sccm and a nitrogen gas having a purity of 99.9% or more flowed at a flow rate of 20 sccm to a source material providing portion 2. Tetraethoxysilane having a purity of 99.5% or more contained in a bubbler 4 was heated to 70°C. Nitrogen gas flowed at a flow rate of 95 sccm and hydrogen chloride gas diluted to 5% with nitrogen flowed at a flow rate of 20 sccm to the source material providing portion 2. Nitrogen gas flowed at a flow rate of 280 sccm to the source material providing portion 2. The total flow rate was adjusted to 800 sccm by adding diluted nitrogen gas. The total pressure was adjusted to 1 atm. The oxygen partial pressure was 3.3×10^{-1} atm, and the tetraethoxysilane partial pressure was 3.3×10^{-3} atm. A sapphire (0001) (nonconformity in lattices of 3.3%) substrate was heated to a temperature within a range of 550 to 850°C. A quartz epitaxy thin film having a thickness of 3 μm was formed thereon. The thin film thus obtained was evaluated using double crystal X-ray diffraction (DC-XRD), a scanning electron microscope (SEM), and a reflection infrared spectrometer.

FIG. 2 is a graph showing an X-ray diffraction (XRD) spectrum of the quartz epitaxial film grown at 800°C. Measurement was made with "RIGAKU RINT 2000" under the conditions shown in TABLE 2 below.

TABLE 2

X-ray	Cu _{ka1} 30kV 200mA
Scanning axis	2 θ / θ
Light slit	0.3 mm
Scanning range	35 to 55°

Scanning speed 1°/min

An X-ray diffraction profile in FIG. 2 has a diffraction peak at $2\theta=50.6^\circ$, which shows that the quartz epitaxial film has a hexagonal system (0003). A diffraction peak at $2\theta=41.8$ is a diffraction peak of the sapphire substrate (0006).

The epitaxial film grew in a specific orientation on the surface of the substrate. There is an X-ray diffraction peak representing this specific orientation. According to the present invention, a crystal having a hexagonal system grows on the sapphire substrate (0001), and it is expected to have an X-ray diffraction peak of a quartz (000x). The X-ray diffraction peak obtained represents the hexagonal system quartz phase (0003), which shows that the quartz thin film grew epitaxially in a c-axis orientation.

The FWHM was determined by the double crystal X-ray diffraction peak of the quartz phase (0003) at each growth temperature. FIG. 3 is a plot of the growth temperature and the FWHM. As shown in FIG. 3, the FWHM showing crystallinity decreased as the growth temperature increased. In Example 1, the FWHM of the quartz thin film formed at 850°C was 10.0 minutes. The quartz thin film had excellent crystallinity.

The growth rate of the quartz thin film was calculated based on the film thickness on a sectional photograph of the SEM at each temperature. FIG. 4 is a plot of the growth temperature and the growth rate. Measurement was made using a "SHIMADZU SUPERSCAN-220." Growth of the quartz thin film was started at a temperature of 600°C, and the quartz thin film had a growth rate of 3.0 μm at 850°C. The inverse of the temperature and the growth rate were plotted to calculate the activation energy from the slope thereof. The activation energy was 16.7 kcal/mol. It is reported that a silicon dioxide thin film (insulation film) can generally be formed only at a high temperature of 1000°C or more by thermal decomposition of tetraethoxysilane, or

a reaction of tetraethoxysilane and oxygen. Below 1000°C, SiO₂ will not be deposited on the substrate. The activation energies of the thermal decomposition of tetraethoxysilane, and the reaction of tetraethoxysilane and oxygen are as high as 190 kcal/mol and 230 kcal/mol, respectively. This fact indicates that HCL works catalytically to promote the reaction of the tetraethoxysilane-oxygen-hydrogen chloride-nitrogen system in formation of the quartz thin film.

It was found that a quartz epitaxial thin film having excellent crystallicity could be obtained using the tetraethoxysilane-oxygen-hydrogen chloride-nitrogen system as the silicon source at a lower temperature than that used in a conventional method.

Infrared rays (IR) within a range of 2000 to 4000 cm⁻¹ were irradiated on the quartz epitaxial thin film to measure infrared ray absorption properties. Measurement was made using a "SHIMADZU FTIR-8700." FIG. 5 shows an infrared ray absorption spectrum of the quartz thin film. The quartz thin film was absorbed at a wavenumber of 3585 cm⁻¹, which indicates the presence a hydroxyl group (OH group). The infrared ray absorption spectrum obtained was almost the same as that of a single crystal of quartz. As a result, a quartz epitaxial thin film having excellent crystallicity and optical properties was obtained.

Example 2

Various buffer layers of different thicknesses were formed on the substrate before a quartz epitaxial growth film was formed to determine the effect of the buffer layer. The buffer layer was formed by depositing a quartz at 550°C and annealing it. A quartz epitaxial film was grown thereon. The quartz epitaxial film was evaluated using the DC-XRD to determine the FWHM. The result is shown in FIG. 6. FIG. 6 shows that the crystallicity of the quartz epitaxial thin film changed depending on the thickness of the buffer layer, and the best crystallicity was obtained when the buffer layer had a thickness of about 25 nm. A quartz thin film having

excellent crystallinity absorbed the infrared ray less at a wavenumber of 3585 cm^{-1} indicates the presence of a hydroxyl group (OH group).

According to the present invention, a quartz epitaxial thin film can be provided by a convenient method using a silicon alkoxide such as tetramethoxysilane, tetraethoxysilane, tetrapropoxysilane and tetrabutoxysilane without using a vacuum device. Such a quartz epitaxial thin film is useful for vibrators, oscillators, high frequency filter surface acoustic wave elements, and optical waveguides.

The method for producing the quartz epitaxial thin film of the present invention is an atmospheric pressure vapor phase epitaxial growing method using a silicon alkoxide as a source material. The method of the present invention provides various advantages: (1) the method can be conducted under atmospheric pressure without requiring expensive vacuum device and large-scale equipment, (2) growth rate of a thin film is very high, and production in quantity is possible, (3) no highly purified material is needed, and the thin film can be produced simply and inexpensively, and (4) a thin film with high purity having excellent crystallinity and optical properties can be formed, compared to a conventional laser ablation method or a sputtering method. The method of the present invention is thus highly useful for forming single quartz thin films.

What is claimed is:

1. A quartz thin film made by depositing at least one silicon alkoxide selected from the group consisting of tetramethoxysilane, tetraethoxysilane, tetrapropoxysilane and tetrabutoxysilane on a substrate under atmospheric pressure.
2. A quartz thin film as claimed in claim 1, which is a quartz epitaxial thin film.

ABSTRACT OF THE DISCLOSURE

A single crystal of quartz thin film and a production method therefor are provided. A method for producing a quartz epitaxial thin film comprises the steps of vaporizing a silicon alkoxide as a silicon source under atmospheric pressure to introduce the silicon alkoxide to a substrate with hydrogen chloride as a reaction promoter, and reacting ethyl silicate with oxygen to deposit a quartz on the substrate. The single crystal of quartz thin film has excellent crystallinity, and optical properties.